**Determining the Ionic Formula**

For each ionic compound, draw the positive and negative ions. Add enough ions so the compound has a net zero charge. Write the ionic formula.

1. Mercury (IV) fluoride
2. Lithium oxide
3. Beryllium sulfide
4. Rubidium selenide
5. Sodium chloride
6. Mercury (II) oxide
7. Magnesium iodide
8. Calcium bromide
9. Lead (II) sulfide
10. Sodium iodide
11. Magnesium bromide
12. Lithium sulfide
13. Lead (III)bromide
14. Potassium chloride
15. Mercury(IV) sulfide
16. Iron (II) telluride
17. Potassium oxide
18. Iron (III) iodide
19. Copper (II) chloride